



MACHAKOS UNIVERSITY

University Examinations 2019/2020 Academic year

SCHOOL OF PURE AND APPLIED SCIENCES

DEPARTMENT OF PHYSICAL SCIENCES

FIRST YEAR FIRST SEMESTER EXAMINATION FOR
BACHELOR OF EDUCATION (SPECIAL EDUCATION)
BACHELOR OF SCIENCE (ANALYTICAL CHEMISTRY)
BACHELOR OF SCIENCE (MATHEMATICS)
BACHELOR OF EDUCATION (SCIENCE)
BACHELOR OF SCIENCE (BIOLOGY)

SCH 101: INTRODUCTION TO PHYSICAL CHEMISTRY

DATE: 6/12/2019

TIME: 2.00-4.00 PM

INSTRUCTIONS TO CANDIDATES

- The paper consists of **two** sections.
- Section **A** is **compulsory**.
- Answer any **two** questions from section **B**.

Required data

- Gas constant, $R = 8.314 \text{ JK}^{-1}\text{mol}^{-1} = 0.08206 \text{ L atm K}^{-1} \text{ mol}^{-1}$
- $1 \text{ atm} = 760 \text{ mmHg} = 101325 \text{ Nm}^{-2}$

SECTION A - COMPULSORY.

QUESTION ONE (30 MARKS)

- a) i) State the Gay-Lussacs's law? (1 mark)
- ii) 500.0 liters of a gas in a flexible-walled container are prepared at 700.0 mmHg and 200.0 °C. The gas is placed into a tank under high pressure. When the tank cools to 20.0 °C, the pressure of the gas is 30.0 atm. What is the volume of the cold gas? (3 marks)
- iii) A sample of 0.5 moles of NH_3 gas occupies 0.5 L at 0 °C. Calculate the pressure of the gas (in atm) using the:
- I. Ideal gas equation (2 marks)
- II. Van der Waals equation given $a = 6.58 \text{ L}^2 \text{ atm mol}^{-2}$ and $b = 0.056 \text{ L mol}^{-2}$ (3 marks)
- b) i) State Dalton's Law of partial pressures (1 mark)
- ii) 100 cm^3 of oxygen gas at 1.75 atm and 200 cm^3 of nitrogen gas at 0.6 atm are passed into a 500 cm^3 vessel. Compute the total pressure of the gas mixture in the vessel at the same temperature (4 marks)
- c) i) State the Equilibrium law? (1 mark)
- ii) Given the reversible equation: $\text{N}_{2(\text{g})} + 3\text{H}_{2(\text{g})} \leftrightarrow 2\text{NH}_{3(\text{g})}$ $\Delta H = -5508 \text{ J}$. Explain the effect of the following changes on the position of the equilibrium (4 marks)
- (a) Increase in pressure (b) Reduction in temperature
- (c) Removing ammonia (d) Addition of Hydrogen
- iii) Given equation (i) $\text{N}_{2(\text{g})} + \frac{1}{2}\text{O}_{2(\text{g})} \leftrightarrow \text{N}_2\text{O}_{(\text{g})}$ Whose **K** is 2.0×10^{-2} and equation (ii) $\text{N}_{2(\text{g})} + \text{O}_{2(\text{g})} \leftrightarrow 2\text{NO}_{(\text{g})}$ whose **K** is 3.0×10^{-2} . Compute the **K** for the reaction $2\text{N}_2\text{O}_{(\text{g})} + \text{O}_{2(\text{g})} \leftrightarrow 4\text{NO}_{(\text{g})}$ (4 marks)
- d) The rate of diffusion of O_2 is 0.6litres /sec. Given that the vapour density of gas X is 14 calculate the rate of diffusion for gas X (3 marks)
- e) What is the pH of 0.25M ammonia solution (NH_3) given that $K_b = 1.8 \times 10^{-5}$? (4 marks)

Section B

Answer any two questions

QUESTION TWO (20 MARKS)

- a) Define the following terms;
- Galvanic cells
 - Oxidation number
 - Spontaneous Reaction (3 marks)
- b) Given the following electrochemical cell: $\text{Ag}/\text{Ag}^+(0.5 \text{ M}) // \text{Li}^+(1.5 \text{ M})/\text{Li}$ and the following cell potentials
- $$\text{Li}^+_{(\text{aq})} + \text{e}^- \rightarrow \text{Li} \quad E^\circ = -3.045 \text{ V}$$
- $$\text{Ag}^+_{(\text{aq})} + \text{e}^- \rightarrow \text{Ag} \quad E^\circ = -0.799 \text{ V}$$
- Write the half reactions at the anode, cathode and the overall reaction (3 marks)
 - Determine the e.m.f of the cell at 25 °C. (4 marks)
 - Draw a schematic diagram of the spontaneous cell using separate beakers for each half cell. Label the anode, cathode and the other components of the cell. Indicate the flow of the electrons (5 marks)
- c) i Write the balanced net ionic equation for the reaction
- $$\text{Fe}^{2+} + \text{MnO}_4^- \rightarrow \text{Fe}^{3+} + \text{Mn}^{2+} \quad (3 \text{ marks})$$
- ii Identify the oxidizing and the reducing agent in the above reaction (2 marks)

QUESTION FOUR (20 MARKS)

- a) Define the following terms;
- Rate laws
 - Zero order reaction
 - Rate of reaction (3 marks)
- b) Camphor is composed of 78.90% C and 10.51% O. The remainder is H. What is its empirical formula? (3 marks)
- c) Briefly explain three factors that influence the rate of a chemical reaction (3 marks)
- d) Define half-life ($t_{1/2}$) and show that for a 1st order rate reaction, the half life ($t_{1/2}$) is independent of initial concentration. (5 marks)
- e) The reaction $2\text{NOCl}_{(\text{g})} \rightarrow 2\text{NO}_{(\text{g})} + \text{Cl}_{2(\text{g})}$ is a second order reaction with respect to NOCl. $K=0.81\text{M}^{-1}\text{s}^{-1}$ at 20°C. If the concentration of NOCl is $7.5 \times 10^{-3}\text{M}$, how much NOCl will be left after a reaction time of 10 minutes? Determine the half-life of this reaction (6 marks)

QUESTION FOUR (20 MARKS)

- a) i) State Raoult's Law (1 mark)
- ii) Calculate vapour pressure of water over a solution made by dissolving 225 g of glucose in 575 g of water at 35 °C. (At 35 °C, $P_{\text{H}_2\text{O}} = 42.2$ mmHg, MW of glucose is 180.2 g/mol and H_2O is 18.02 g/mol) (4 marks)
- b) i) Define colligative properties (1 mark)
- ii) List three examples of colligative properties. (3 marks)
- c) Given an equation below for an ideal gas

$$P = \frac{mN\bar{u}^2}{3V}$$

show that the kinetic energy E_k for one mole of a gas is dependent on temperature (3 marks)

- d) Suppose that in the reaction $\text{PBr}_5(\text{g}) \leftrightarrow \text{PBr}_3(\text{g}) + \text{Br}_2(\text{g})$, the fraction of PBr_5 that dissociate to reach equilibrium is γ and the total gas pressure at equilibrium is P . Show that

$$K_p = \frac{\gamma^2 P}{1 - \gamma^2}$$

(5 marks)

- e) Given that K_{sp} for AgBr is $2.8 \times 10^{-10} \text{ mol}^2\text{l}^{-2}$ at 25 °C. Calculate the solubility of AgBr in 1 litre of water. (3 marks)

QUESTION FIVE (20 MARKS)

- a) i) What is an ideal solution? (1 mark)
- ii) Briefly describe the properties of an ideal solution. (2 marks)
- b) Calculate the vapour pressure of a solution containing 5.85 g benzene (MW = 78) and 2.3 g of toluene (MW = 92) at 323K, if the vapour pressures of the pure components at this temperature are $3.6 \times 10^4 \text{ Nm}^{-2}$ and $1.12 \times 10^4 \text{ Nm}^{-2}$, respectively. (4 marks)
- c) A chemist isolated a bacterial gene fragment and dissolved a 17.6 mg sample of the material in enough water to make 31.5 mL of solution. The osmotic pressure of the solution was 0.340 torr at 25.0 °C. Calculate the molar mass of the gene fragment. (4 marks)
- d) i) What is a buffer solution? (1 mark)
- ii) Derive the Henderson-Hasselbalch equation for a weak acid (CH_3COOH) (4 marks)
- e) Sketch the following graphs:
- i) Pressure against volume based on Boyle's law (2 marks)
- ii) Volume against temperature based on Charles's law (2 marks)